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KHAPILOV, Yu.; mladshiy nauchnyy sotrudnik; ZHURILOV, V., mladshiy nauchnyy sotrudnik

Use by foreign countries of plastics and synthetic materials in shipbuilding (from "Quarterly Transactions of the Institute of the Institute of Mayal Architecture," no.3, July 1958). Mor.flot 19 no.8:

38-40 Ag '59, (MIRA 12:11)

1. Institut komplekanykh transportnykh problem AN SSSR, (Shipbuilding)

SYRMAY, A.G., nauchnyy sotr.; OBERMEYSTER, A.M., nauchnyy sotr.;
HRONFMAN, A.I., nauchnyy sotr.; SHIMKO, K.N., kand. tekhn.
nauk; PARAKHONSKIY, B.M., kand. ekon. nauk. Prinimali uchastiye: ZHURILOV. V.I., nauchnyy sotr.; ZUBKOV, M.I., nauchnyy
sotr.; SHVARTS, G.L., nauchnyy sotr.; MIKHEYEV, A.P., doktor
tekhn. nauk, prof., otv. red.; EYKOV, I.K., red. izd-va;
DOROKHINA, I., tekhn. red.

[Water and air transportation in capitalist countries: trends in the development of equipment] Vodnyi i vozdushnyi transport kapitalisticheskikh stran; tendentsii razvitiia tekhnicheskikh sredstv. Moskva, Izd-vo Akad.nauk SSSR, 1961. 350 p. (MIRA 15:1)

1. Akademiya nauk SSSR. Institut kompleksnykh transportnykh problem.

(Merchant marine) (Aeronautics, Commercial)

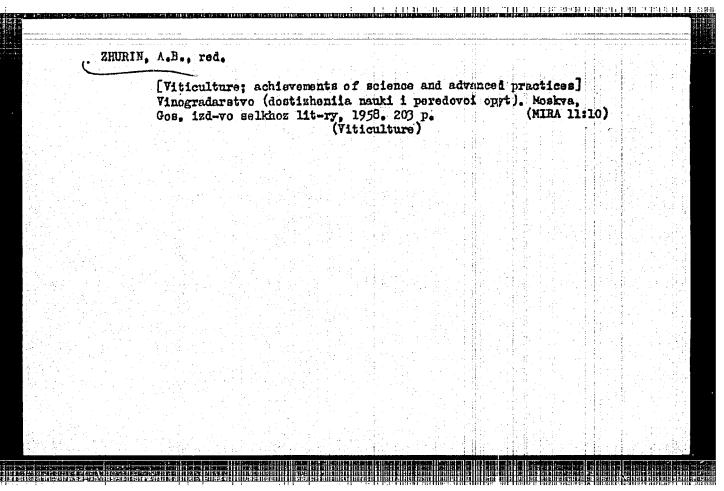
SYRMAY, A.G. Prinimali uchastiye. ZHURILOV, V.I., mlad. naushnyy sotr.;
KANTOROVICH, Ya.B., kand. tekhm. nauk, retsenzent; VORONOV, Ye.K.,glav.
ekonomist, retsenzent; OBERMEYSTER, A.M., otv. red.; DOBSHITS, K.L.,
red. izd-va; SUSHKOVA, L.A., tekhm. red.

[Method of deciding upon the running speed andcarrying capacity of
seagoing vessels] Metodika obosnovaniia skorosti knoda i gruzopod[®]
emnosti morskikh sudov. Moskva, Izd-vo Akad. nauk SSSR, 1961. 50 p.
(MIRA 14:11)

1. Gosudarstvennyy proyektno-konstruktorskiy i nauchno-issledovatel'skiy institut morskogo transporta Ministerstva morskogo flota SSSR
(for Voronov).2. Institut kompleksnykh transportnykh problem AN SSSR
(for Zhurilov).

(Naval architecture)

ZHURIN, A.			4:		*****		· · · · · · · · · · · · · · · · · · ·
"Hydromechanization of the pla	inting of vi	neyards on	the koll	cohozes	and sta	te	
farms in the Soviet Union."							
p. 143 (Mezhduna Rodnyi Selsko	okohoziaist v	ennyi Zhurr	nal, Vol.	2, No.	2, 195	8,	
Sofia, Bulgaria).							
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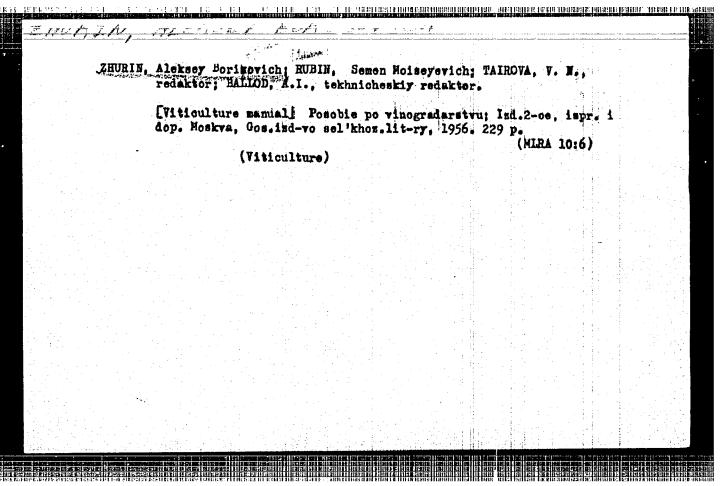


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GUSEVA, A.M.; SHEFFER, V.V.; SHIN, P.V.; ZHURIN, A.B.; TIKHONOV, N.P.;
KLYUSHKIN, P.A.; PUL'SON. R.Kh.

Local information. Zashch. rast. ot vred. 1 bol. 8
no.10:59-60 0'63. (MIRA 17:6)

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			(Phyl	loxera)			



Rubin, S. M. A manual on viticulture Moskva, Gos, izd-vo sel'khoz. Lit-ry, 1950. 222p. 1. Viticulture—Handbooks, manuals, etc. 2. Vituulture—Russia. I. Zhurin, A. B. jt. au.											
A manual on viticulture Moskva, Gos, izd-vo sel'khoz. Lit-ry, 1950. 222p. 1. Viticulture-Handbooks, manuals, etc. 2. Vituulture-Russia. I. Zhurin, A. B.	IURIN, A. B., Ju. au.					:			. : -		
A manual on viticulture Moskva, Gos, izd-vo sel'khoz. Lit-ry, 1950. 222p. 1. Viticulture-Handbooks, manuals, etc. 2. Vituulture -Russia. I. Zhurin, A. B.	Rubin. S. M.										
1. Viticulture-Handbooks, manuals, etc. 2. VitaultureRussia. I. Zhurin, A. B.	A manual on viticulture Moskva, Gos	3, 12d-V	ro səl	L'khoz.	Lit-	ry, 19	50.	222p.			
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USSR/Cultivated Plants = Fruits. Berries.

Abs Jour : Ref Zhur = Biol., No 7, 1958, 30086

Author : Zakharova, Ye.I., Zhurin, A.B.

Inst : Title : The Viticulture of Hungary.

Orig Pub : Sad i ogorod, 1957, No 9, 60-63.

Abstract : No abstract.

KOLESNIKOV, Venedikt Andreyevich, prof., doktor sel'skokhoz.neuk; ZHURIH.

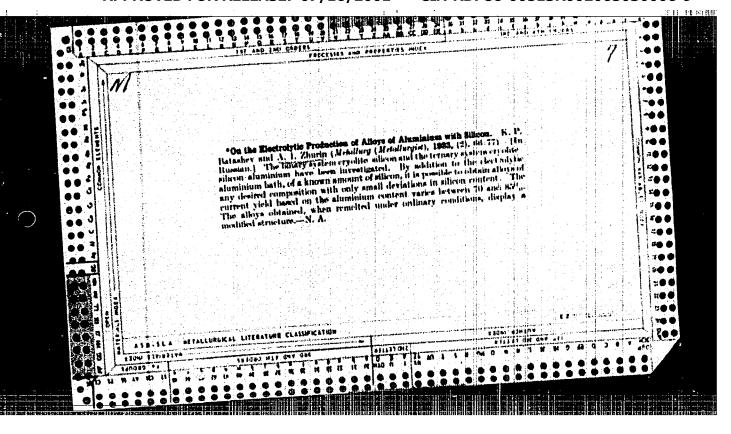
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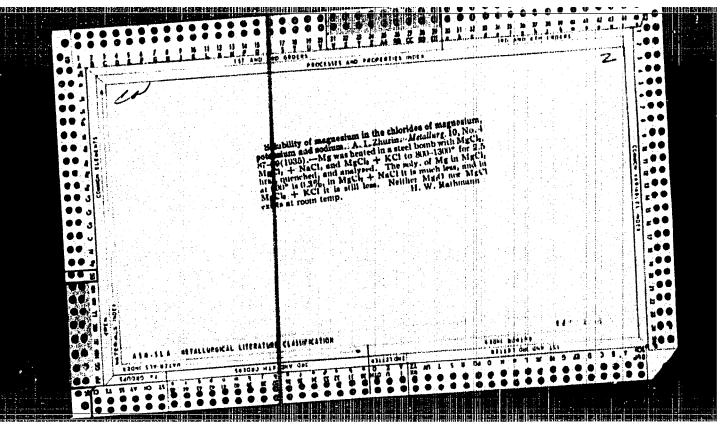
"gronom; KAPTSINZL'

Anna Petrovna, agronom; KOVAL', Alia Alekseyevich, sentomolog; ZUBAREV, N.A.; LUR'TH, B.D., red.; RIZGULYAYEVA, N.G., tekhn.red.

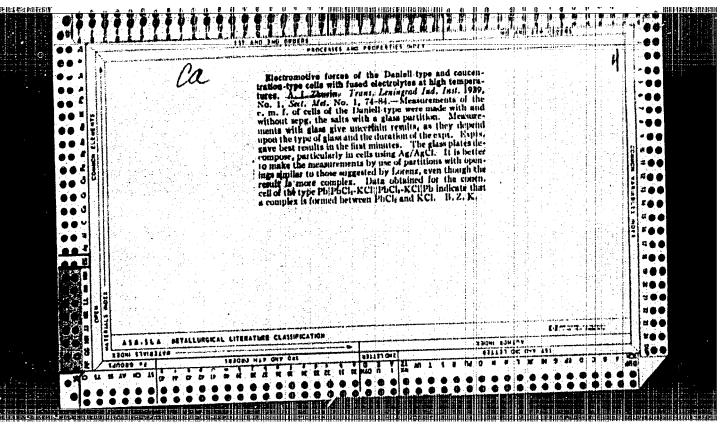
[Amateur fruitgrower's reference menual] Kalendar'-spravochnik sadovoda-liubitelia. Moskva, Izd-vo M-va sel'.khoz.SSSR, 1959.

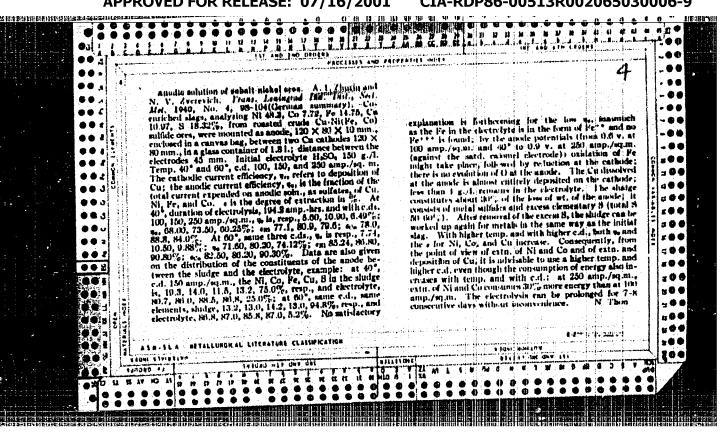
Mode p. (Fruit culture)

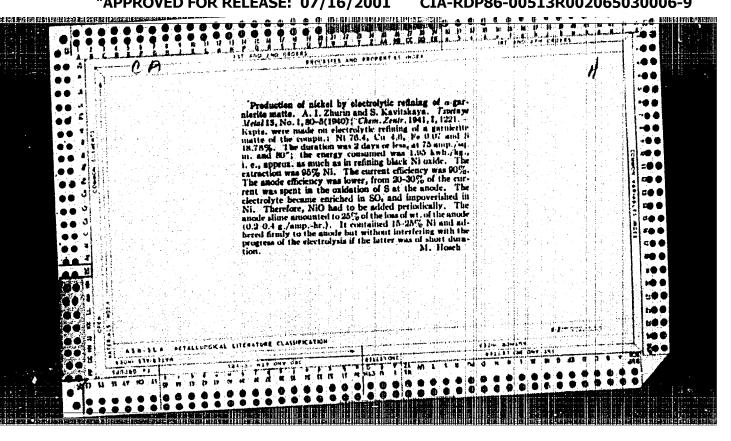


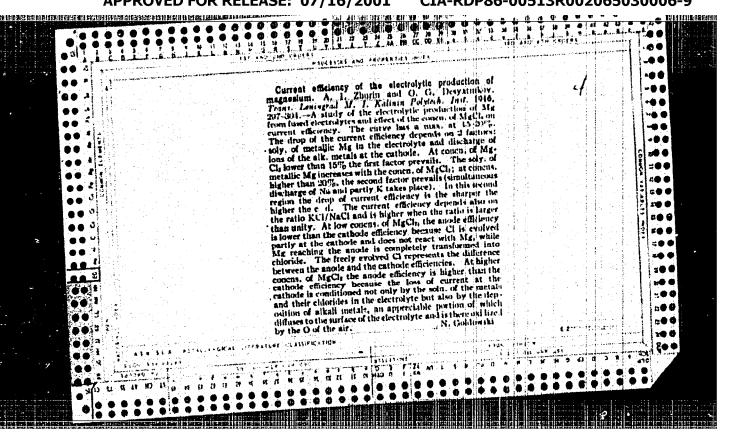


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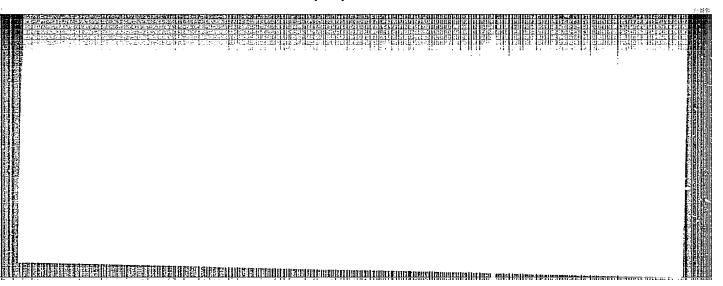








ZHURIN, A	.I.;S	ноукнит,	M.G														
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137-58-6-11979

TO THE PERSON OF THE PERSON OF

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 6, p 113 (USSR)

AUTHORS: Zhurin, A.I., Shoykhet, M.G.

TITLE:

-- Properties

Buffering Properties of Nickel Sulfate Solutions and the Formation of Hydrates in These Solutions (O bufernykh svoystvakh rastvorov sul'fata nikelya i gidratoobrazovaniya v nikh)

PERIODICAL: Tr. Leningr. politekhn. in-ta, 1957, Nr 188, pp 173-180

ABSTRACT: ... The incipient formation of hydrates in Ni electrolytes was investigated experimentally. Some considerations are presented concerning the discrepancy between the pH data on the formation of hydrates as given by A.L. Rotinyan and V.Ya. Zel'des (Zh. prikl. khimii, 1950, Vol 23, p 717) and the data obtained in earlier research on this problem. In addition, the authors comment on the mechanism of the action of such buffer additives as H₃BO₃, (NH₄)₂SO₄, and CH₃COOH in the course of the electrolysis. See also RzhMet, 1957, Nr 4, abstract 5717. 1. Electrolytes--Properties 2. Nickel sulfate solutions

3. Hydrates--Analysis

Card 1/1

Zhurin, A.I.

137-58-5-9307

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 5, p 74 (USSR)

AUTHORS

Zhurin, A.L., Shoykhet, M.G.

<u>ra di conservación de que a se servición en la discontración en en entración de contractor de contractor de co</u>

TITLE:

The Effect of Organic-compound Additives on the Process of Electrolytic Deposition of Nickel From Sulfate Solutions (Vliyaniye primesey organicheskikh soyedineniy na elektroliticheskoye osazhdeniye nikelya iz sul'fatnykh rastvorov)

PERIODICAL: Tr. Leningr. politekhn. in-ta, 1957, Nr 188, pp 181-190

ABSTRACT:

A study of the effect of certain organic compounds on the current efficiency and the quality of metal being deposited during electrolytic refining of Ni. It is established that of all compounds which are leached out of wood by the electrolyte, the watersoluble constituents of wood and linen rag are the most harmful. On conversion to C content, the content of water-soluble compounds must not exceed 20 mg/l. As the solution is freed from Fe and Co, the organic compounds become oxidized and are removed. Whenever large amounts of wood or linen rag are introduced into the process, it is essential that they be treated preliminarily with hot water for a period of 1-2 days so as to remove water-soluble compounds contained in the surface layer. Wood may be treated with a 2% lye solution.

Card 1/1

1. Nickel--Electrodeposition 2. Electrolytes--Properties 3. Electrolysis--Effectiveness 4. Organic compounds--Flectrolysis

Zhurin, A.I.

137-58-5-9306

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Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 5, p 74 (USSR)

AUTHORS: Zhurin, A.I., Ivanov, L.A.

TITLE: Electrolytic Precipitation of Nickel From Suifale Solutions With

Addition of Ammonium Salts (Elektroliticheskoye osazhdeniye nikelya iz sul'fatnykh rastvorov s primeneniyem dobavok ammo-

niynykh soley)

PERIODICAL: Tr. Leningr. politekhn. in-ta, 1957, Nr 188, pp 191-203

ABSTRACT: Studies were performed in order to determine conditions

most suitable for the precipitation of Ni from solutions containing buffering additives in the form of ammonium salts; the quality of the Ni precipitates was also studied. It was established that range of the buffer action of solutions buffered with ammonium salts is greater than that of solutions buffered with boric acid. Good-quality elastic deposits are obtained from sulfate solutions buffered with ammonium sulfate containing small amounts of Cl ion (5 g/l). The S and H content in these deposits

is not greater than in deposits obtained from solutions with

boric acid.

Card 1/1

1. Nickel--Electrodeposition 2. Ammonium salts--Applications

3. Electrolytes--Properties

137-58-6-12028

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 6, p 120 (USSR)

Zhurin, A.I., Pyunnenen, S.P. AUTHOR:

Combined Influence of Additions of Manganese with Iron, TITLE:

Cobalt, and Antimony Present in Solutions During Electrolytic Deposition of Zinc (Sovmestnoye vliyaniye primesi margantsa s zhelezom, kobal'tom i sur'moy v rastvorakh pri elektroliti-

cheskom osazhdenii tsinka)

Tr. Leningr. politekhn. in-ta, 1957, Nr 188, pp 204-211 PERIODICAL:

The influence of individual admixtures (Mn, Fe, Co, Sb) as ABSTRACT:

well as the combined effect of Mn and Fe, Mn and C, Mn and Sb, and Sb and Co were studied in the process of electrolytic deposition of Zn. The following facts were established: 1) the presence of a considerable quantity of Mn2+ ions (up to 5-20 g/f) results in a marked reduction in current yield; this condition is due to the oxidation of Mn²⁺ to MnO4 and the reduction of MnO4 to Mn2+ ; 2) compared with Mn the Fe reduces the current yield even more abruptly; this is explained by the fact that the Fe²⁺ is oxidized to Fe³⁺ and that the Fe may be

Card 1/2

137-58-6-12028

Combined Influence of Additions of Manganese (cont.)

deposited on the cathode (accompanied by intense evolution of H2); introducing gelatin increases the current yield. 3) When simultaneously present in the same solution, the elements Mn and Fe mutually reduce each other's action, a fact which is attributable to mutual oxidation-reduction processes occurring in the electrolyte (E); 4) Combined action of Mn and Co, Mn and Sb, and Co and Sb reduces the current yield to a greater degree than could be expected in the case of concurrent but independent action; this condition is explained by the assumption that the more abrupt change in the surface of the cathode (as compared with the action of only a single ingredient) is responsible for a more abrupt change in the density of current; 5) introduction of gelatin into the E greatly suppresses the action of the impurities, particularly of such substances as Sb, Co, etc., i.e., impurities which are separated out at the cathode; 6) experiments in which a baffle was employed have shown that the current yield of Zn in a neutral E is very great even if considerable quantities of impurities are present.

1. Zinc-Electrolytic deposition 2. Electrolytes-Chemical properties 3. Manganese-Chemical reactions 4. Iron-Chemical reactions 5. Only

3. Manganese--Chemical reactions 4. Iron--Chemical reactions 5. Cobalt--Chemical reactions 6. Antimony--Chemical reactions

Card 2/2

ZHURIN, A.I.

137-58-5-10269

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 5, p 198 (USSR)

AUTHORS: Gvozdeva, I. I., Zhurin, A. I.

TITLE: The Electrochemical Properties of Rhenium (Elektrokhimiches-

kiye svoystva reniya)

PERIODICAL: Tr. Leningr. politekhn. in-ta, 1957, Nr 188, pp 212-224

ABSTRACT: The following questions are investigated in this study: a) the relationship of Re plating potential to the strength of the KReO4 (5, 10, 15, 20 g KReO4/liter; 60°C, pH 1.3; b) the effect of pH on plating potential (60°, 10 g KReO4/liter, pH varied from 0.7 on plating potential (60°, 10 g KReO4/liter, pH varied from 0.7 to 1.4); c) effect of temperature (pH 1.3, 10 g/liter, Cu cathode, temperature varied from 20 to 90°); d) effect of cathode material (Re, Ni, Mo, Cu, Fe, 10 g KReO4/liter, pH 1.3, 60°). It is found that the equilibrium potential of Re in a solution containing 10 g KReO4/liter and 15 g H2SO4/liter is +0.353 v at 30°. The balance of electrolysis products on the electrodes indicates the occurrence of a process of O2 liberation at the anode, while two processes - liberation of Re and of H2 - occur at the cathode.

The optimum conditions for deposition of Re from a sulfate Card 1/2. KReO4 solution are determined. The best Re coatings at

Th	e Electroche	emical Proj	perties of	Rhenium			137-58-5.	
ma obt 0.9	ximum curr ained when -1), at 15 ar	ent efficien l liter of wa mps/dm ² an	cy and mir ater containd 85-900.	nimum co ns 15 g K	nsumpti KReO ₄ an	on of elected l2-15 g	ctric pow 3 H ₂ SO ₄ (er are pH
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AUTHOR:

Zhurin, A.I.

On Electrolytic Nickel Refining in Sulfate-Chloride and

TITLE:

Chlorous Solutions

PERIODICAL:

Izvestiya vysshikh uchebnykh zavedeniy, Tavetnaya metallurgiya,

1960, No 3, pp 54 - 61

TEXT: Information is given on results of experimental investigations into processes of electrolytic Ni refining in sulfate-chloride and pure chlorous solutions. The experiments were performed with the participation of M.G. Shoykhet, V.S. Ponomarev, B.P. Gorshkov, and A.F. Vikharev. The following processes were studied: the effect of chlorine ions on anode and cathode processes were studied: the effect of chlorine ions on anode and cathode potentials; the effect of Cl concentration on current efficiency; the effect of the circulation rate and current density on current efficiency and the of the circulation rate and current density on current efficiency and the centration of the cathode nickel; long-lasting electrolysis with different conquality of the cathode nickel; long-lasting electrolysis with different concentration of chlorine ions; behavior of metals of the platinum group; recentration of chlorous nickel solutions from iron and cobalt. The totality of fining of chlorous nickel solutions from iron and cobalt. The totality of results obtained leads to the following conclusions; pure chlorous solutions

Card 1/2

X 1812 P74 E A CONTRACTOR OF THE STREET OF On Electrolytic Nickel Refining in Sulfate-Chloride and Chlorous Solutions with high concentration of Ni ions or sulfate-chloride solutions with high concentration of Ni and chlorine with a relatively low content of sulfate ions may be used to intensify the operation of electrolytic shops; there are no considerable investment costs required; the efficiency of the shops would increase by a factor of 1.5 - 2. Moreover, the efficiency of the electrolytic bath may be raised by using high concentrations of Ni, raised current densities and reduced circulation rates of the solution, without enlarging the capacity of devices for the purification from iron and cobalt. There are 6 tables, 6 graphs and 5 references: 3 Soviet and 2 English. ASSOCIATION: Leningradskiy politekhnicheskiy institut (Leningrad Polytechnical Institut), Kafedra elektropirometallurgii tsvetnykh metallov (The Chair of Electropyrometallurgy of Non-Ferrous Metals) SUBMITTED: June 6, 1969 Card 2/2

BAYMAKOV, Yuriy Vladimirovich: ZHURIN, Aleksandr Ivanovich: LEVIN, A.I., prof., doktor tekhn. nauk, retsenzent; SMIRNOV, V.I., prof., retsenzent; STENDER, V.V., prof., retsenzent; CORBUNOVA, K.M., prof., doktor khim. nauk, red.; PAKHOMOVA, G.N., kand. tekhn. nauk, red.; MARENKOV, Ye.A., red.; MISHARINA, E.D., red.izd-va; MIKHAYLOVA, V.V., tekhn. red.

[Electrolysis in hydrometallurgy]Elektroliz v gidrometallurgii. Moskva, Metallurgizdat, 1963. 616 p. (MIRA 16:2)

1. Kafedra tekhnologii elektrokhimicheskikh proizvodstv Ural'skogo politekhnicheskogo instituta (for Levin). 2. Kafedra metallurgii tsvetnykh metallov Ural'skogo politekhnicheskogo instituta, Deystvitel'nyy chlen Akademii nauk Kazakhskoy SSR (for Smirnov).

3. Chlen-korrespondent Akademii nauk Kazakhskoy SSR (for Stender).

(Hydrometallurgy) (Electrometallurgy)

ACCESSION NR: AT4026277

B/2563/63/000/223/0069/0074

AUTHOR: Zhurin, A.I.; Li, Hang-kuan

HISTORIAN STOLING THE SECTION OF THE PROPERTY OF THE PROPERTY

TITLE: Electrolytic purification of crude indium containing tin and cadmium

SOURCE: Leningrad. Politekhnicheskiy institut. Trudy*, no. 223, 1963. Metallurgiya tsvetny*kh metallov (Metallurgy of nonferrous metals), 69-74

TOPIC TAGS: indium, indium refining, electrolytic refining, indium purification, indium electrolysis, tin, cadmium.

ABSTRACT: Crude indium may be purified by many methods, including electrolysis. In the present investigation, the authors used electrolytic purification of indium on solid indium anodes and Ti cathodes. The crude indium contained tin and cadmium. Spectral snalysis was used to determine the content of tin and cadmium in the material. The accuracy of measurement was generally 0.001% and in some cases was increased to 0.0005%. It was found that during the anode dissolution of crude indium, most of the tin remains in the sludge, while cadmium together with indium pass into the solution (see Fig. 1 of the Enel.). Tin is deposited together with indium on the cathode, only its extraction coefficient is less than unity and depends inversely on the current density. Cadmium is deposited together with indium only at high concentrations. Complete elimination of tin and cadmium from solution indium only at high concentrations.

	may be accomplished by hydrogen sulfide precipitation in two stages at pH 0.37 and then 0.9. At pH=2, the current yield at the cathode is essentially 100% and that at the anode exceeds is a chloride. "The authors express gratitude to Ye. L. Grinzayd, Associate Professor of quantitative analysis of tin and cadmium." Orig. art. has: 3 figures and 2 tables.
	ASSOCIATION: Leningradskiy politekhnicheskiy institut (Leningrad Polytechnic Institute) SUBMITTED: 00 DATE ACQ: 16Apr64 ENCL: 01 SUB CODE: MM NO REF SOV: 003 OTHER: 001
c	ard 2/31

Electrolytic refining of crude indium containing tin and cadmium. Trudy IPI no.223:69-74 '63. (MIRA 17:11)	
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ACCESSION NR: AT4026278

8/2563/63/000/223/0075/0081

AUTHOR: Zhurin, A.I.,; Chao, Ching-sheng

TITLE: The direct extraction of gallium from aluminate solutions by electrolysis

SOURCE: Leningrad. Politekhnichoskiy institut. Trudy*, no. 223, 1963. Metallurgiya tsveiny*kh metallov (Metallurgy of nonferrous metals), 75-81

TOPIC TAGS: gallium electrolysis, aluminum, aluminum refining, gallium, alluminate solution, electrolysis, mercury cathode, gallium extraction

ABSTRACT: Gallium accompanies aluminum in all its ores, including bauxite, but is normally lost during refining. The authors previously proposed the electrolytic extraction of gallium from aluminate solutions in one stage using a mercury cathode. In the present paper, they study the possibility of extracting gallium from aluminate solutions in one step using a solid cathode. In clarifying the optimal conditions for this procedure, they determined the potentials for the extraction of hydrogen from alkaline solutions on steel and gallium cathodes, as well as the extraction of gallium and hydrogen from an alkaline solution of sodium gallate. Finally, experiments were carried out on gallium extraction from both sodium gallate and aluminate solutions. The results show that it is theoretically possible

Card 1/2

ACCESSION NR: AT402627	8	
urrents of $400-600 \text{ a/m}^2$.	allium from aluminate solutions containectrodes. Electrolysis should be perfeasing steel cathodes previously coated bined with the production of hydrogen s	ormed at 60-80C and
SSOCIATION: Leningrade	ky politekhnicheskiy institut (Leningra	d Polytechnic Institute)
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JB CODE: MM	NO REF SOV: 003	OTHER: 002
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Direct Trudy	recover	ry of 223:75	gallin -85	um from	alum	inate	sol	utions	by	electronic (MIRA	lysis. 17:11)	
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ACCESSION NR: AT4026279

8/2563/63/000/223/0082/0086

AUTHOR: Zhurin, A. I.; Ovchinnikov, A. V.

TITLE: Some of the electrochemical properties of indium

SOURCE: Leningrad. Politekhnicheskiy institut. Trudy*, no. 223, 1963. Metallurgiya tsvetny*kh metallov (Metallurgy of nonferrous metals), 82-86

TOPIC TAGS: indium, anode polarization, cathode polarization, indium electrochemistry, electrochemistry

ABSTRACT: Indiam is acquiring great importance among the rare elements, but little has been published on its electrochemistry. For this reason, the authors measured the anode and cathode polarization curves for indiam in solutions of its chloride, the overvoltage required to evolve H₂, and the yield at the anode and cathode during electrolytic refining of indiam. It was found that the anodic dissolution of indiam in a 0.407 N solution of its chloride proceeds at a high rate with little polarization at potentials from -0.45 to -0.42 volts, monovalent and trivalent ions being produced simultaneously. The proportion of monovalent ions increases with the current density. Meanwhile, deposition of indiam at the cathode also takes place at a rapid rate with little polarization. When the current

1/2

density reached a maximum value, simultaneous discharge of indium lons and hydrogen ions takes place at the cathode. The lower the pH of the solution, the lower this maximal current density. Orig. art. has: 3 figures, 5 chemical formulas, and 2 tables. ASSOCIATION: Leningradskiy politekhnicheskiy institut (Leningrad Polytechnic Institute) SUBMITTED: 00 DATE ACQ: 16Apr64 ENCL: 00 SUB CODE: MM, GC NO REF SOV: 001 OTHER: 004				14 14 1 14 14 14 14 14 14 14 14 14 14 14	
density reached a maximum value, simultaneous discharge of indium lons and hydrogen ions takes place at the cathode. The lower the pH of the solution, the lower this maximal the lower this maximal and 2 tables. ASSOCIATION: Leningradskiy politekhnicheskiy institut (Leningrad Polytechnic Institute) SUBMITTED: 00 DATE ACQ: 16Apr64 SUB CODE: MM, GC NO REF SOV: 001 OTHER: 004	ACCESSION NR: AT40262	179			
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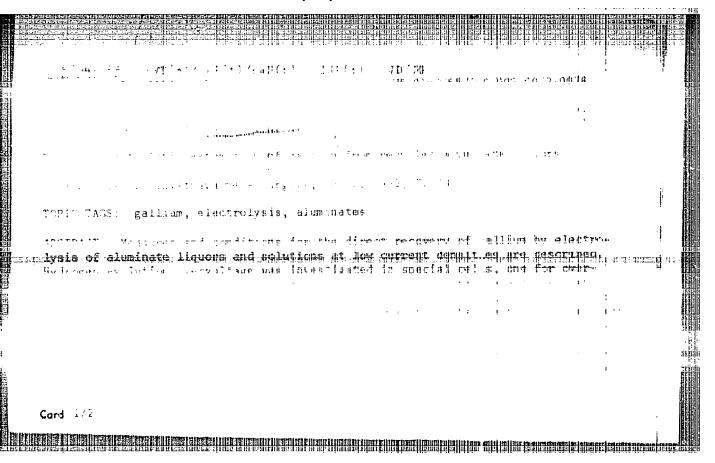
Gallium recovery from return aluminate alkalies by electrolysis.

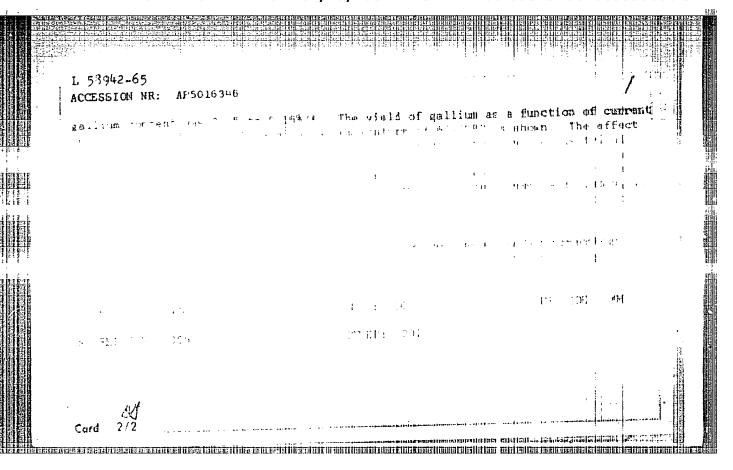
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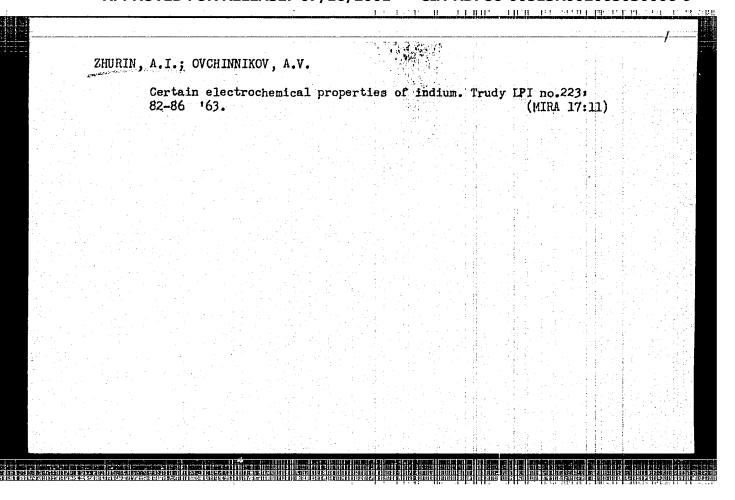
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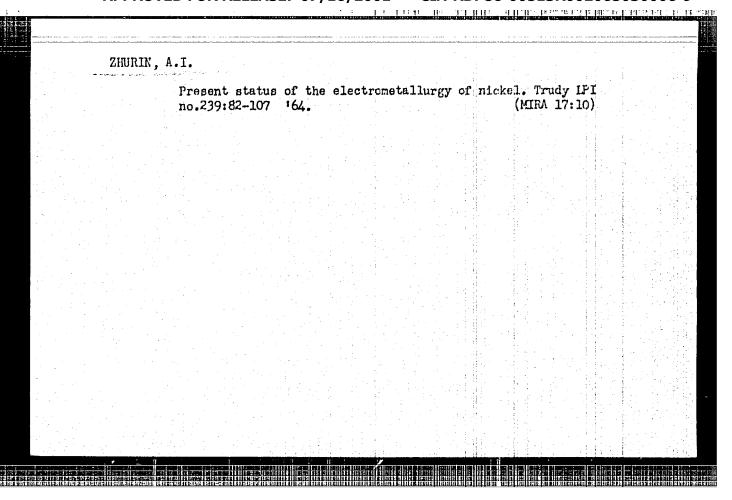
politekhnicheskogo instituta. Submitted December 24, 1963.

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AYZENSHTADT, L.A.; PEN'KOV, P.M.; GLADKOV, B.A.; LIKHT, L.O.;

KRIMMER, T.Ye.; KASHEPAV, M.Ya., kand. tekhh. næuk;

MERPERT, M.P., kand. tekhn. næuk; KOPERBAKH, B.L.;

CHERNIKOV, S.S., kand. tekhn.nauk; BELOV, V.S.; ZHURIN.

B.F.; MONAKHOV, G.A., kand. tekhn.nauk; MORCTOV, I.I.;

MUSHTAYEV, A.F.; OGNEV, N.N.; PALEY, M.B., kand. tekhn.

næuk; FURMAN, D.B.; LIVSHITS, A.L., kand. tekhn.næuk; MECHETNER,

B.Kh.; SOSENKO, A.B; AVDULOV, A.N.; LEVIN, A.A., kand. tekhn.

næuk; YAKOBSON, M.O., doktor tekhn.nauk; MAYOROVA, E.A.,

kand. tekhn.nauk; MOROZOVA, Ye.M.; ZUSMAN, V.G., kand. tekhn.

næuk; NAYDIS, V.A., kand. tekhn.næuk; VIADZIYEVSKIY, A.P., prof.,

doktor tekhn. nauk, red.; BELOGUR-YASNOVSKAYA, R.I., red.;

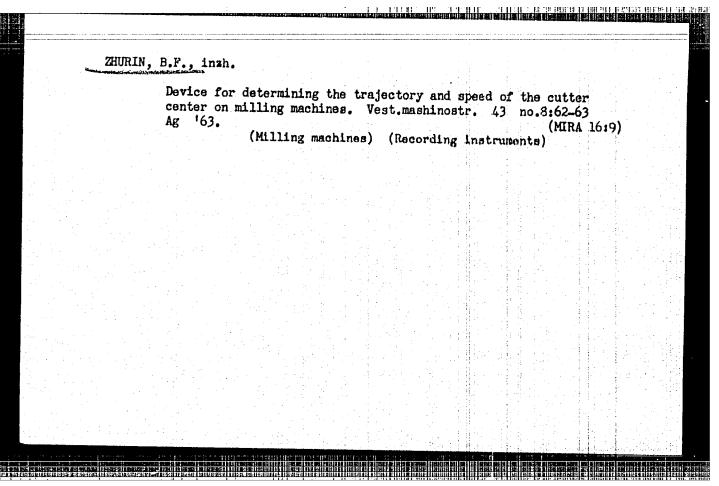
CHIGAREVA, E.I., red.; ASVAL'DOV, M.Ya., red.; KOGAN, F.L.,

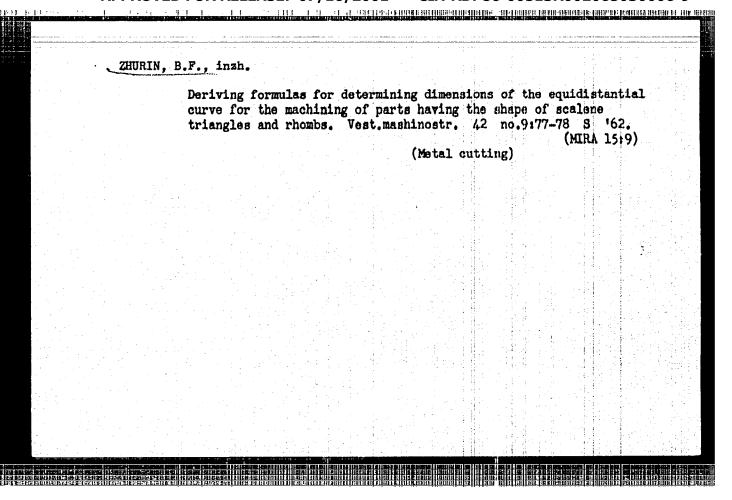
tekhn. red.

[Machine-tool industry in capitalist countries] Stankostroenie v kapitalisticheskikh stranakh. Pod red. i s predisl. A.P.Vladzievskogo. Moskva, 1962. 822 p. (MIRA 15:7)

1. Moscow. TSentral'nyy institut nauchno-tekhnicheskoy informatsii mashinostroyeniya. 2. Eksperimental'nyy nauchnoissledovatel'skiy institut metallorezhushchikh stankov (for Vladziyevskiy, Belogur-Yasnovskaya, Chigareva, Asval'dov, Kogan).

(Machine-tool industry)





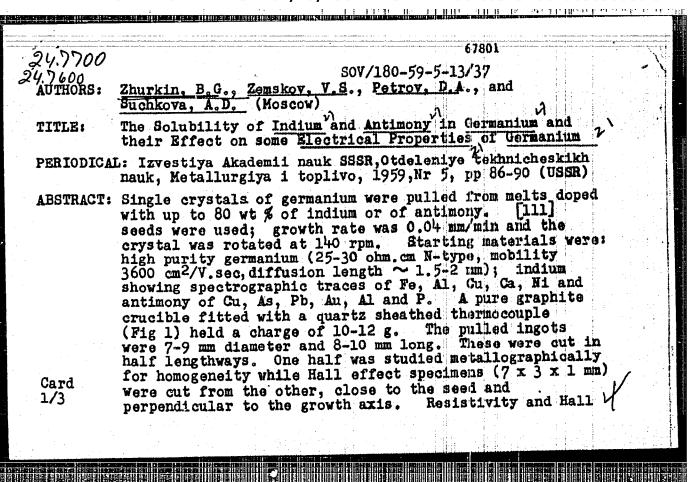
t .	Zhurin, Boris Ivanovich		11/5 831.1	
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	Roditel'skaya Obshchestvennost' v Promoshch	'Shkole		
	(The Parents in the Community as an Aid for	the Schools)		
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	Moskva, Uchpedgiz, 1955			
	157 P. Illus., Tables.			
	Bibliographical Footnotes.			
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المناسب والم 67293 24.7700 sov/180-59-4-26/48 Zhurkin, B.G., Zemskov, V.S., Petrov, D.A. and **AUTHORS:** (Moscow) Suchkova, A.D. The Nature of the Quasi-Binary Germanium-Indium-Antimony TITLE: System PERIODICAL: Izvestiya Akademii nauk SSSR, Otdeleniye tekhnicheskikh nauk, Metallurgiya i toplivo, 1959, Nr 4, pp 156-158 (USSR) Germanium with electron-type conduction and a specific ABSTRACT: resistance 25 to 30 ohm/cm was used together with zone refined antimony and indium. Crystals were pulled from the melt. Results are given in Table 1. All the samples had electron-type conductivity and samples with high insb content had a higher concentration of electrons than those The number of current carriers with low InSb content. to $1.9 \times 10^{19}/\text{cm}^3$. The value for varied from 1.2×10^{18} fully compensated additions is 2.5 x 1013/cm3. Thus there was an excess of Sb atoms. Experiments were carried out using the same Ge:Sb ratio and increasing the In content. Results are given in Table 2. With a ratio of Inish of 2.5 there is still electronic conduction very near to the compensated alloy. With In:Sb = 4.4 there is hole-type Microstructures were examined along the conduction. Card 1/2

The Nature of the Quasi-Binary Germanium-Indium-Antimony System
length of the crystal pulled from a melt. A second phase appears (see Fig) which from microhardness tests temperatures greater than 650°C - the temperature of dissociation of Insb. The authors conclude that because not possess the properties of a quasi-binary system are Soviet and 2 English.

SUBMITTED: March 16, 1959

Card 2/2



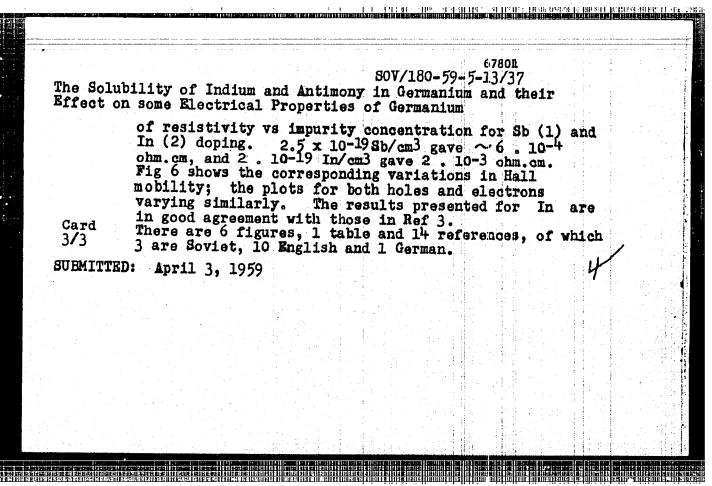
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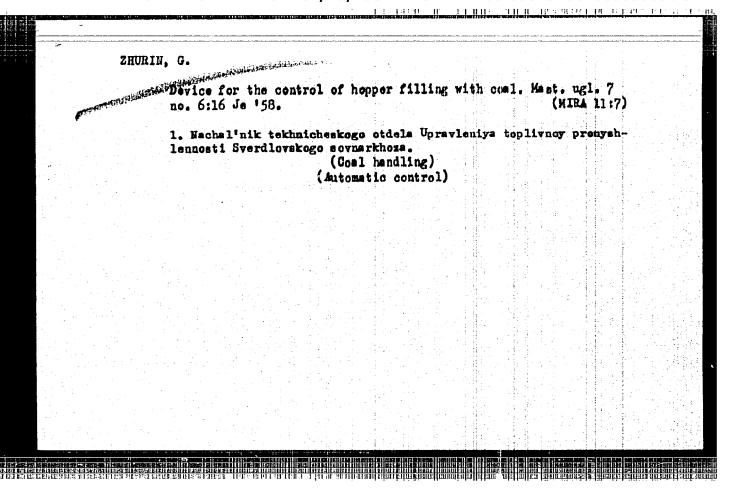
sov/180-59-5-13/37

The Solubility of Indium and Antimony in Germanium and their Effect on some Electrical Properties of Germanium

emf were measured with a potentiometer type PPTN-1 and a galvanometer type M-25/3. Resistivity measurements were ± 5% but Hall measurements (3700 Oe field) for the higher impurity concentrations had greater errors, from 10-50%. In determining impurity concentrations from resistivity and Hall measurements complete ionization and degeneracy were assumed. The table shows equilibrium concentrations of indium and antimony in solid and liquid germanium at various temperatures (both wt % and at % values are given). The corresponding phase diagrams are plotted in Figs 3 and 4 (compositions in at %). Solid Ge containing 6.6.10-2 at % In is in equilibrium with a melt containing 71.6 at % In at 620 oC, and solid germanium containing 7.2 . 10-2 at % Sb with liquid containing 70.5 at % Sb at 693 oC. Extrapolation to the eutectic horizontals suggests maximum solid solubilities of 8 . 10-2 at % In and about 0.1 at % Sb. No retrograde solid solubility was found for Sb. Fig 5 shows log-log plots (which are linear)

Card 2/3





GORODETSKIY, David Yevseyevich; ZHURIN, Origoriy Mikhaylovich;

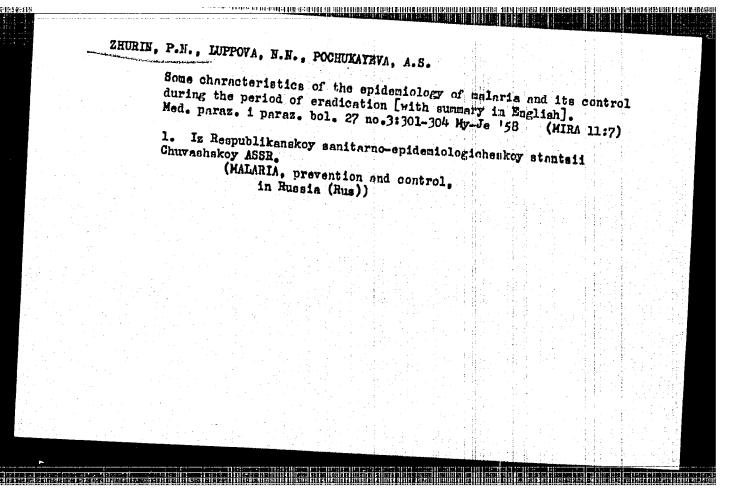
ZUBAREV, Leonid Aleksandrovich; ADAMOVA, L., red.;
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[Put the reserves of the fuel industry to use]Rezervy toplivnoi promyshlemosti v deistvii. Sverdlovsk, Sverdlovskoe
knizhnoe ixl-vo, 1961. 110 p. (MIRA 15:8)

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KUKINOV,	V.M.; MASOKIN, V.I.; ZHURIN, N. Ya.; RODZEVILLO, I.T.
	New equipment and progressive technology. Bezop. truda v prom. 8 no. 9:31-33 S *64 (MIRA 18:1)
	1. Nachal'nik Gubkinskoy rayonnoy gornotekhmicheskoy inspektsii (for Kukinov). 2. Shakhta imeni Gubkina (for Masckin, Zhurin, Rodzevillo).



ZHURIN, P.N., dotsent

Intracellular inclusions in spring catarrh. Vest.oft. 69 no.2:9-11
Mr-Ap '56.

1. Is Chuvashskogo nauchno-issledovatel'skogo trakhomatownogo instituta
(dir.--dotsent P.A.Shishkin; nauchnyy rukovoditel'--dotsent Ts.Yu.

Komenetskaya)

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USSR / Virology. Viruses of Man and Animals. Chlamydozoa.

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Abs Jour

: Ref Zhur - Biologiya, No 22, 1958, No. 99185

Author

: Zhurin, P. N.

Inst

: State Scientific Research Institute for Eye Diseases

Title

: Diagnostic Meaning of Preparation - Imprints From

Mucosa of the Eyelids

Orig Pub

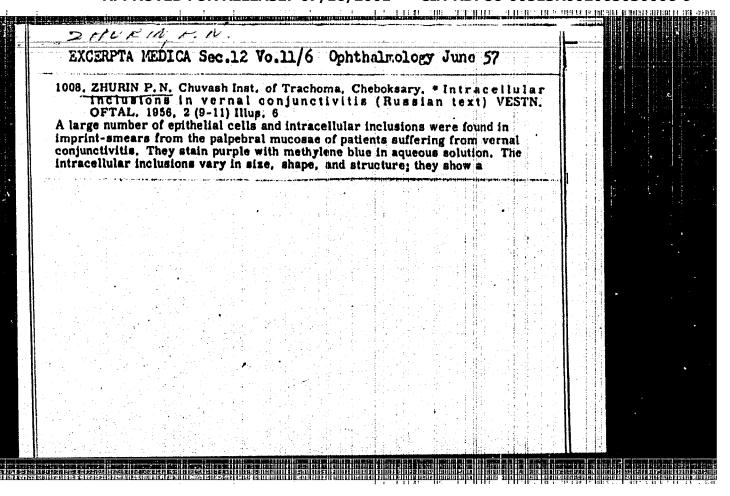
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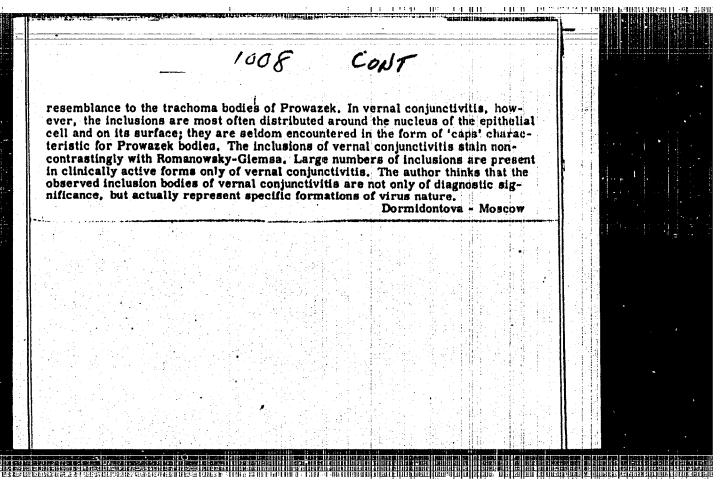
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Abstract : No abstract given

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Diagnosis of New Trachoma Infections based on the presence of Prowazek's Excise Vest. Oftalmol., 28, No. 6, 1949. Chuvash Sci. Res. Trachoma Inst., -c1949	
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USSR/Virology. Chlamydozon.

Abs Jour: Ref Zhur-Biol., No 17, 1958, 76538.

Author: Zhurin, P. N.

Inst

Title: Morphology and Dynamics of Changes of Prowazek's
Intra-Cellular Bodies in the Course of an Infectious

Orig Pub: Sb. nauchn. tr. Chuvashk. n.-i. trachoratozn. in-t, 1957, vyp. 2, 166-176.

Process with Trachora.

Abstract: Serial scrapings of the conjunctiva epithelium of 72 patients with trachorn were studied. It is proposed to divide the Prowazek's bodies found in different stages of the disease according to form into compact, friable (three degrees) and diffuses and according to size into small,

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USSR/Virology. Chlamydozoa.

Abs Jour: Ref Zhur-Biol., No 17, 1958, 76538.

trachorn, it is proposed to take into account the norphological peculiarities of the Prowazek bodies.

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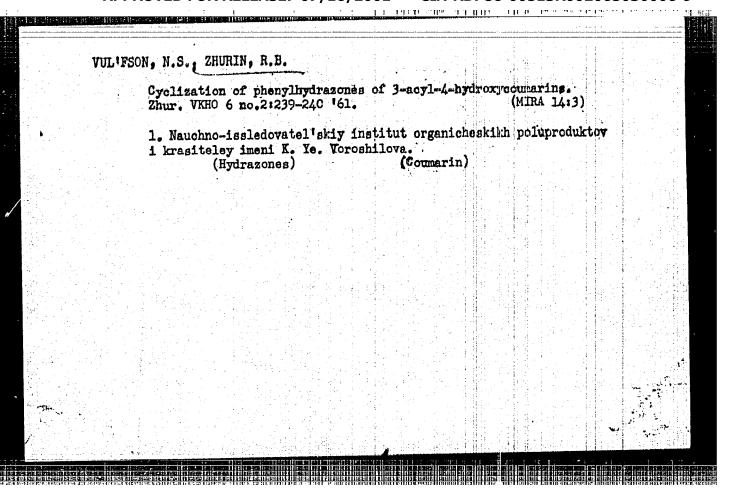
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ZHURIN,	R.B.; VULIFSON, N.S.
	Reaction of C-acyla ion of heterocyclic ketoenols. Part 2: Synthesis of X-acotyl- and X-propionyl- Y-phenyltetronic acids. Zhur.ob.khim. 30 no.8:2467-2468 Ag 160. (MIRA 13:8)
	l. Nauchno-issledovatel'skiy institut organicheskikh poluproduktov i krasiteley imeni K.Ye. Voroshilova. (Tetronic acid)

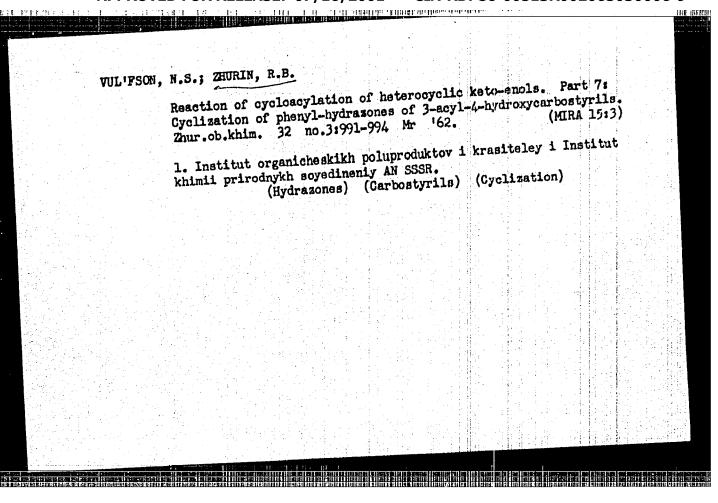
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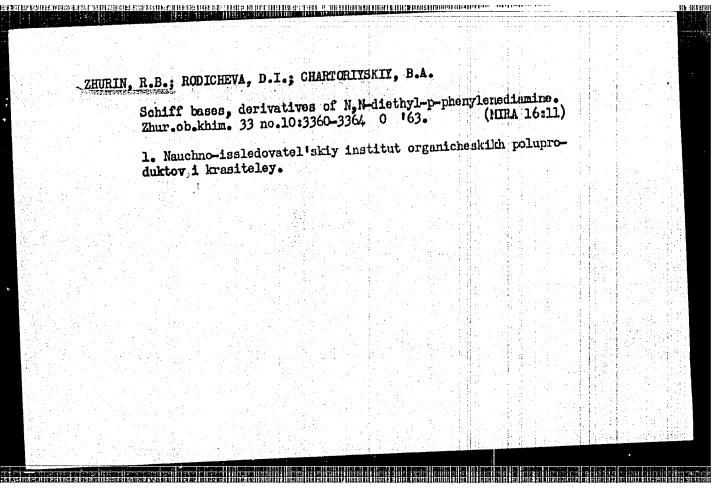
R.B.; VUL'FSON, N.S. C-Acylation of heterocyclic the C-acylation of 4-hydroxy no.3:875-879 Mr. 161.	ketoenols. Par coumarin. Zhur	t 5: Mechanism of ob. khim. 31 (MIRA 14	.:3)
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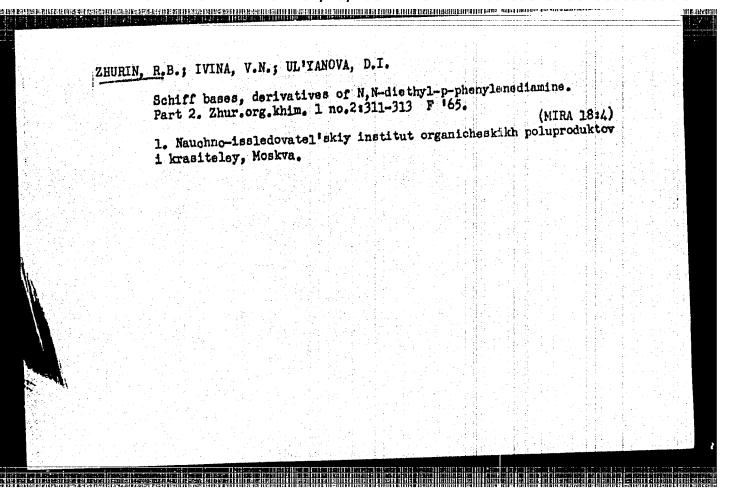
31 no.10:3381-3385 0 '61. (MIRA 14:10) 1. Institut organicheskikh poluproduktov i krasiteley imeni K.Ye.Voroshilova. (Couparin) (Hydrazones)	C-acylation of heterocycl of phenylhydrazones of 3-	acyr-4-nyuroxycoumarin	E - ZHUL OD	ion .khim. RA 14:10)
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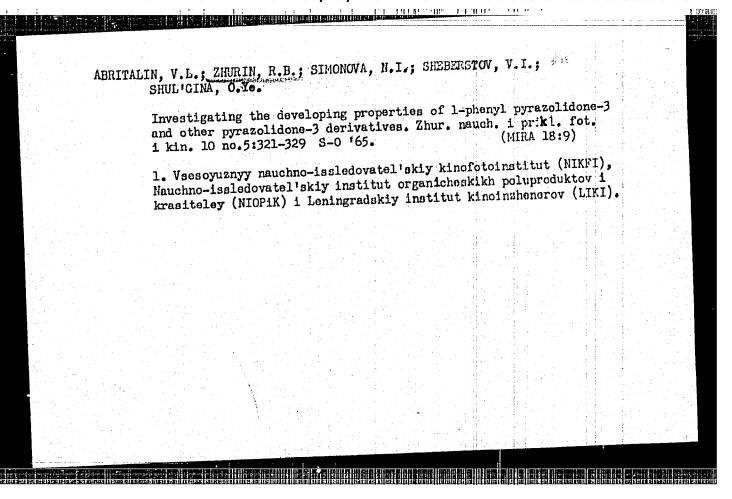


	Some derivatives of 3-pyrazolidinone. Zhur.ob.khim. 31 no.8:2758-2761 Ag '61. (MIRA 14:8)
	1. Nauchno-issldeovatel'skiy institut organicheskikh polu- produktov i krasiteley imeni K. Ye. Voroshilova; Nauchno- issledovatel'skiy kino-fotoinstitut i Leningradskiy institut kinoinzhenerov.
	(Pyrazolidinone)

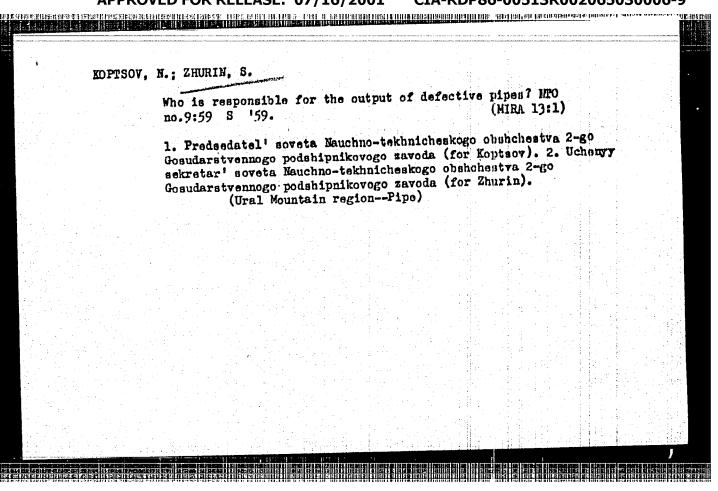






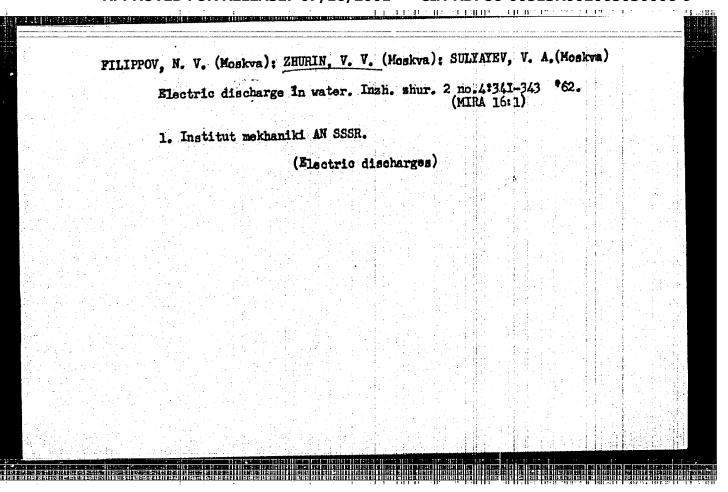


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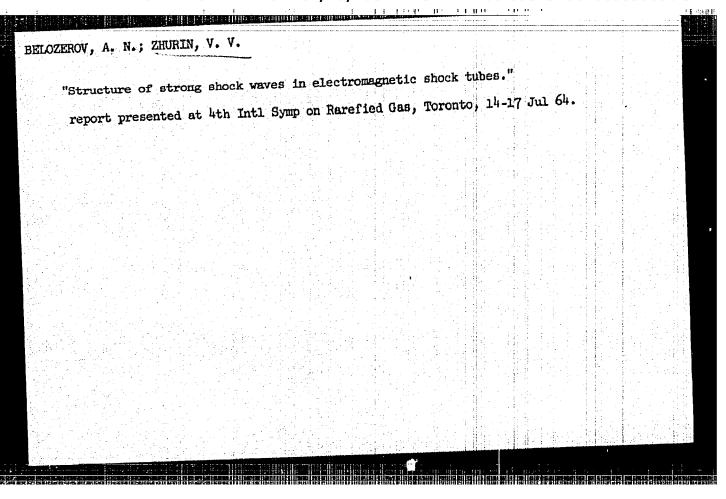
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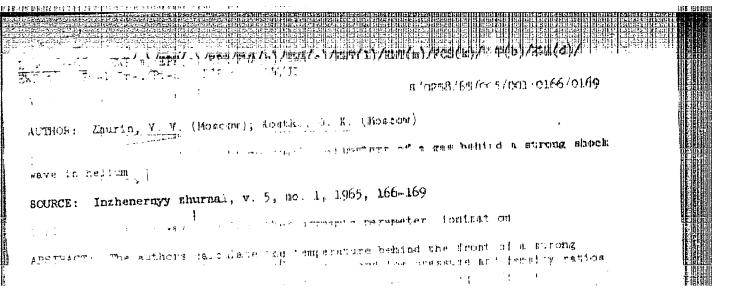


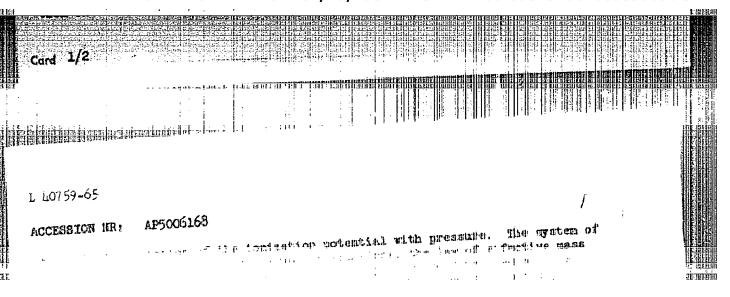
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ACCESSION NR: AP3000724 \$/0258/63/003/002/0373/0375	
AUTHORS: Zhurin, V. V. (Moscow); Sulynyav, V. A. (Moscow); Hukoviskiy, V. M. (Moscow)	3)
TITLE: Shock waves in electromagnetic shock tube SOURCE: Inzhenernyy zhurnal, v. 3, no. 2, 1963, 373-375 TOPIC TAGS: shock wave, ionization, plasma, magnetic dipole, shock tube,	
ABSTRACT: The technique of obtaining strong shocks at high ionization levels in electromagnetic shock tubes was studied. The discharge was obtained from a bank of capacitors (18 microfarad capacity) discharging at 20-kv. Quantz tubes 1000 in length and with internal diameters of 11 and 40 mm acted as slock tubes. The gases used were helium and hydrogen, with an initial pressure between 0.05-5 mm Hg. Oscilloscopes and high-speed movie cameras were used to record data. In the	wa
experiment pressures up to 2000 atm. were obtained with shock speeds of 8 x 107 cm/sec. The electron gas temperature behind the shock was estimated at 700 ev. Unlike the observation by F. R. Scott and R. F. Wenzel (Phys. Fluids. vol. 2, No. 6, 1959) no magnetic dipole was observed in the ionized gas. However, ds. Card 1/2	. ESPENISI

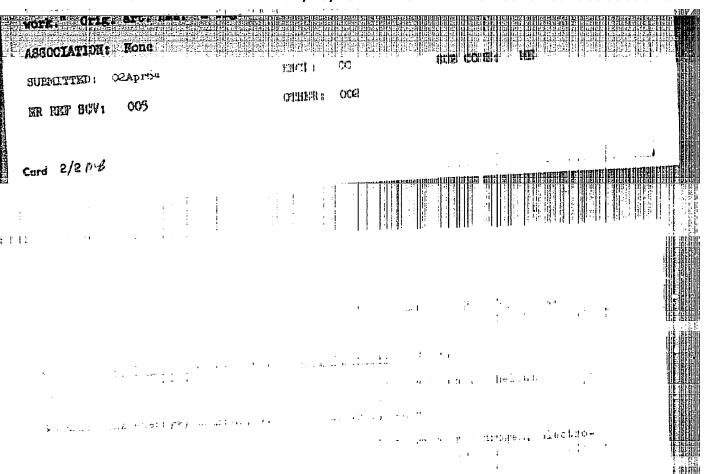
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previous investigators ha	author is grateful to A.	A. NIKOLISKIV TOT	UIB INTERBUL
in this investigation and his valuable discussions	to N. V. Filippov. S. R	. Kholovi and Al.	A LABORCY TOT
ASSOCIATION: Institut me	khanik AN SSSR (Institut	e of Mechanica, A	SSR)
SUBMITTED: 12Sep62	DATE ACQ: 21Jun6	3	BNGL: 00
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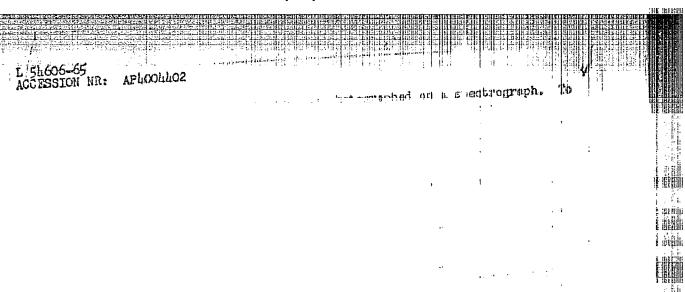


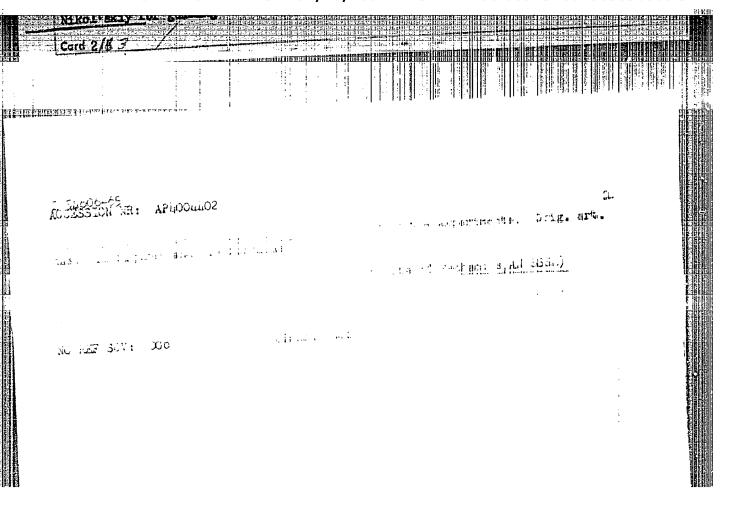
"On some phenomena associated with propagation of magnetically driven strong shock waves through gases".	
report presented at the 2nd All-Union Congress on Theoretical and Applied Mechanics, Moscow, 29 Jan - 5 Feb 64.	
and Applied Mechanics, Moscon, 27	
하는 사람들은 그릇을 모두 많은 그를 받는 것 같아. 그는 그를 모르게 그릇을 걸려가 하를 모르는 것	
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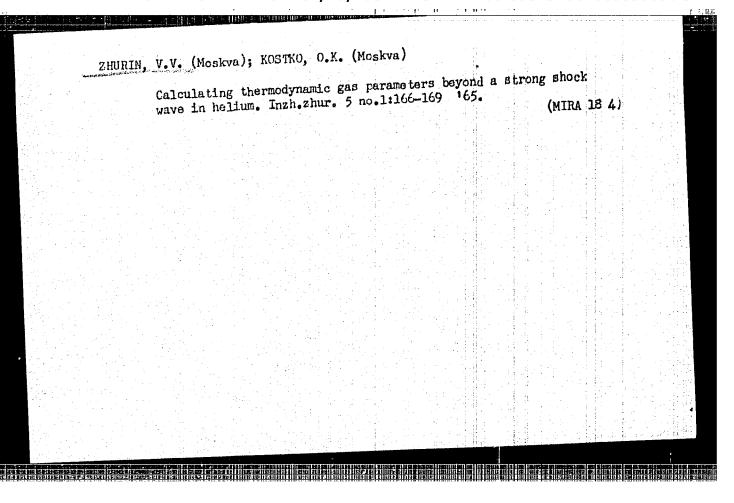












VUL'FSON, N.S.; ZHURINA, F.G.; SENYAVINA, L.B.

Reformatorskii reaction with bromomalonic ester. Fart 3: Further study of the reaction of bromomalonic ester with benzaldehyde. Zhur. ob. khim. 34 no.7:2344-3347 11 *64 (MERA 17:8)

l. Institut khimii prirodnykh soyedineniy AN SSSR i Nauchmoissledovatel skiy institut organicheskikh poluproduktov i krasiteley.

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AUTHORS:

Rozental', L. V., Burdygina, G. I., Korneva, E. D.,

Zhurina, F. G.

TITLE:

Plasticization of Triacetate Cellulose Films by Means of

Ester Mixtures of Higher Synthetic Fatty Acids

PERIODICAL:

Khimicheskaya promyshlennosti, 1960, No. 5, pp. 15 - 18

This paper deals with a study of the plasticizing effect of esters of higher synthetic fatty acids (C6 - C16). It follows from the experiments that low temperatures favor the combination of the plasticizer with triacetate cellulose even when all traces of diluents and solvents are removed from the film. At higher temperatures and higher relative atmospheric humidity, this combination is checked. The number of double bendings endured by the film increases with the increase in the number of carbon atoms in the alcohol radical of fatty acid ester; under the same conditions the plasticizing effect also increases at lower temperatures. For the same number of carbon atoms in the acid

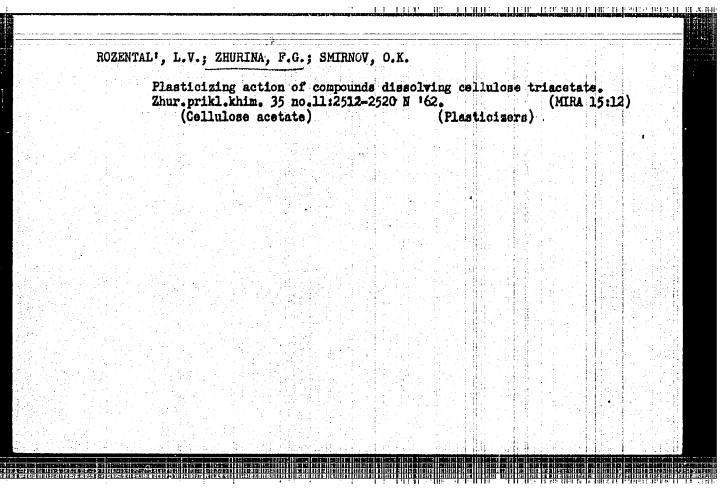
Card 1/2

Plasticization of Triacetate Cellulose
Films by Means of Ester Mixtures of
Higher Synthetic Fatty Acids

radical of the esters of higher synthetic fatty acids; their plasticizing effect increases with the increase in the molecular weight of the alcohol radical. There are 5 tables and 4 references: 3 Soviet and 1 German.

ASSOCIATION: NIKFI (Motion Picture and Photography Scientific Research Institute). NIOPik im. K. Ye. Voroshilova (Scientific Research Institute of Organic Semifinished Materials and Dyes imeni K. Ye. Voroshilov)

Card 2/2



ROZEFFAL', L.v.; BURDYGINA, G.I.; KORNEVA, E.D.; ZHURDIA, F.G.

Plasticization of cellulose triacetate films by mixtures of esters of higher synthetic fatty acids. Khim.prom. no.5:367-370 J1-Ag'60.

(MIRA 13:9)

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(Cellulose acetate) (Plasticizers) (Acids, Fatty)